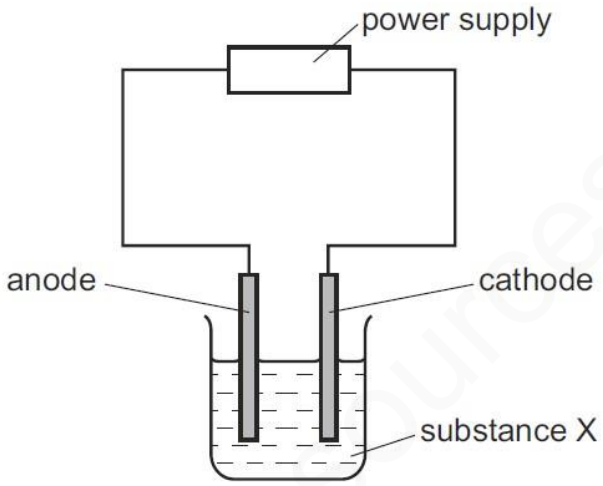
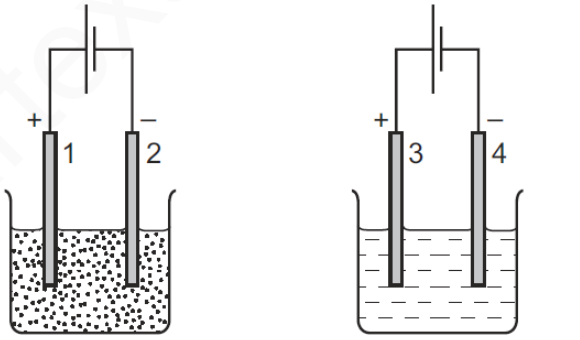
NO:	**ELECTROLYSIS OF MOLTEN IONIC COMPOUNDS-SET-2**
1	Substance X was electrolysed in an electrolytic cell. A coloured gas was formed at the anode and a metal was formed at the cathode.  What is substance X? **A** aqueous sodium chloride **B** molten lead bromide **C** molten zinc oxide **D** solid sodium chloride

MS-1	B
2	Metals could be extracted from their molten chlorides using electrolysis. Which substances are formed at each electrode? **A** chlorine / hydrogen **B** chlorine / metal **C** hydrogen / metal **D** metal / chlorine
MS-2	B
3	Two electrolysis experiments were carried out as shown in the diagram below. The graphite electrodes are labelled 1-4.  molten sodium chloride concentrated aqueous sodium chloride Which row describes the products at the electrodes in these experiments?

	electrode 1	electrode 2	electrode 3	electrode 4
A	chlorine	hydrogen	chlorine	hydrogen
B	chlorine	sodium	chlorine	hydrogen
C	chlorine	sodium	hydrogen	chlorine
D	sodium	chlorine	sodium	chlorine

MS-3	B
4	I One molten compound and two aqueous solutions were electrolysed. The table gives the compounds electrolysed and the electrodes used. In which experiments is a gas evolved at the cathode? **A** 1, 2 and 3 **B** 1 and 2 only **C** 1 only **D** 3 only
MS-4	B

	substance electrolysed	electrodes
1	concentrated hydrochloric acid	carbon
2	concentrated sodium chloride	platinum
3	molten lead bromide	platinum

5	Which row describes the electrolysis of molten potassium bromide? **A** bromine — hydrogen **B** bromine — potassium **C** hydrogen — bromine **D** potassium — bromine
MS-5	B